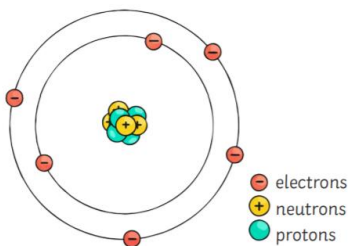


Draw and label an atom. Include labels for the following: neutron, proton, electron.

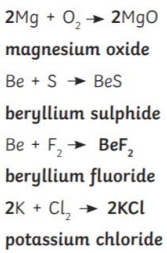


True or false?  
 1. The radius of an atom is 0.1nm **True**  
 2. Most of the mass is in the shell of the atom. **False, most of the mass is in the centre**

What are the symbols for the following elements?

Element	Symbol
oxygen	O
lithium	Li
sodium	Na
potassium	K
helium	He
carbon	C
magnesium	Mg

Complete and balance the following equations. What is the name of the compound formed?



Mixtures  
 Write the definition of a mixture. Give two examples.

**Two or more elements together, not chemically joined and can be easily separated.**

**Salt water, sand and water**

Name the compounds and the elements they contain.

- NaCl - **sodium chloride, sodium and chlorine**
- MgO - **magnesium oxide, magnesium and oxygen**
- MgS - **magnesium sulfide, magnesium and sulfur**
- FeS - **iron sulfide, iron and sulfur**

What is the ratio of the elements in the following compounds?

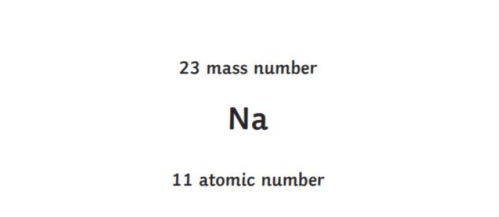
- e.g. CaO = **1:1**                      NaCl = **1:1**
- MgCl<sub>2</sub> = **1:2**                          lithium fluoride = **1:1**
- K<sub>2</sub>O = **2:1**                                sodium hydroxide = **1:1:1**

Fill in the table to show the charges and mass of the components of an atom.

Name	Charge	Relative Mass
proton	+1	1
neutron	0	1
electron	-1	very small

What is the overall charge of an atom?  
**No charge**

Complete the following diagram for sodium, include the atomic number and the atomic mass number.



What is the mass number?  
**Total number of protons and neutrons.**  
 How do you calculate neutron number?  
**Atomic mass – proton number**

**Isotopes** are elements with a different number of neutrons but the same number of protons, e.g. carbon 12 and carbon 14.

How can you use isotopes to calculate the relative atomic mass? Write down the equation.

$A_r = \frac{\text{sum of (isotope abundance} \times \text{isotope mass number)}}{\text{sum of abundances of all the isotopes.}}$

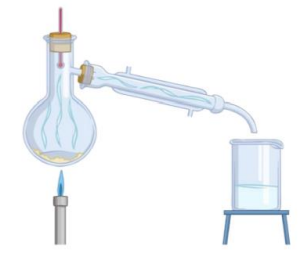
A compound is 2 or more **elements**, chemically **joined**.

Which of the following are compounds?  
 Put a ring round them.

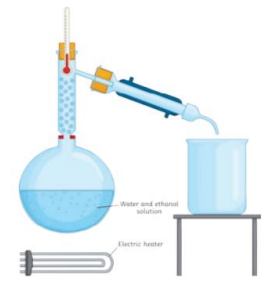
oxygen, salt water, **magnesium oxide, sodium chloride, nitrogen**

Why have you circled the ones you have?  
**They have 2 or more elements in the word equation.**

**Separating Mixtures**  
 What are the following separation techniques?



**Distillation**



**Fractional distillation**

What separation technique would you use to separate out different inks in pens?

**Chromatography**

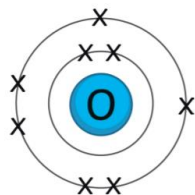
How can salt be collected using the process of crystallisation?  
**By heating up a mixture of salt and water, the water will evaporate and leave the salt in the bowl.**

Sand and water can be separated by using a process called **filtration**.

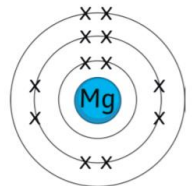
Describe, in 4 steps, how to collect salt from rock salt.

- 1. Grind the mixture;**
- 2. Add water and stir;**
- 3. Filter the mixture;**
- 4. Evaporate the salt water and salt is left over.**

Complete the electronic structure diagrams for:  
oxygen



magnesium



Describe why the noble gases are so unreactive.

**Their outer shell is full of electrons.**

The boiling points of the noble gases **increase** as you go down the group.

**This is because there are more forces to bond the atoms together, therefore more energy is required to break the bonds.**

Describe what happens to the reactivity of the alkali metals as you go down the group.

**It increases**

Why?

**The number of electrons increases. They are further away from the nucleus. There is less pull on the outer electrons so the atom is more likely to lose an electron.**

Complete the word and symbol equation for sodium reacting with water:

sodium + water → sodium hydroxide + hydrogen



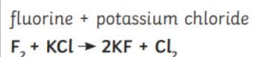
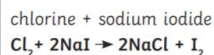
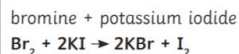
List 3 halogens  
**chlorine, fluorine, iodine, astatine**

How many electrons do they have in their outer shell?  
**7 electrons**

Describe how the reactivity changes as you go down the group.

**They become less reactive, the atom becomes larger because there are more electron shells, further from the nucleus so the pull of the nucleus is less. So the electron is less likely to be gained as there is less of a positive pull.**

Write balanced symbol equations for the following reactions:



Underline the properties of metals and circle the properties of non-metals:

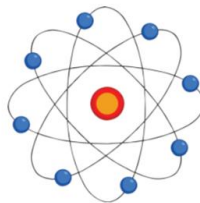
Strong, low density, malleable, dull, good conductors of heat and electricity, high melting and boiling point, brittle, not good conductors of electricity

James Chadwick discovered the...  
(underline the correct answer)

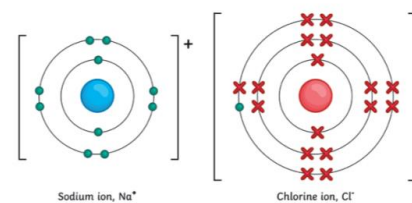
proton

**neutron**

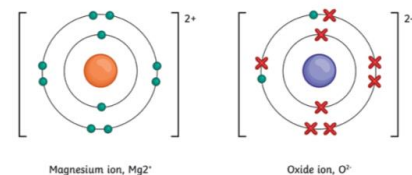
electron



Complete the following dot and cross diagrams for:  
**NaCl**



**MgO**



Complete word equations for the following reactions:

sodium + chlorine → **sodium chloride**

lithium + iodine → **lithium iodide**

potassium + bromine → **potassium bromide**

How are the groups arranged in the periodic table?

**According to their properties.**

How can you tell that the alkali metals are very reactive?

**According to their properties.**

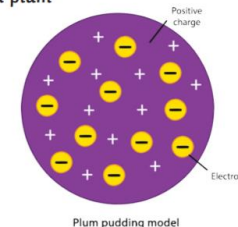
How can you tell the noble gases are unreactive?

**Full shell of outer electrons.**

Describe the plum pudding model of the atom.

Draw a diagram.

**A sphere of positive charge with electrons dotted about; looking like a plum pudding.**



Why did scientists believe this model?  
**Lack of experimental evidence.**

Describe what the alpha scattering experiment showed scientists.

**Most alpha particles go straight through, some are scattered, some rebound off the gold foil. This shows that the nucleus of an atom has a very small radius. Most of the mass is concentrated in the nucleus.**

Niels Bohr discovered that

**electrons orbit the nucleus in shells.**

Why did Mendeleev leave gaps in the periodic table?

**He knew that the elements existed but they hadn't been found, based on their mass.**

What happened to some of the gaps he left?

**They have been filled. Scientists have found some of the elements.**